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## Chemical Formulas And Chemical

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Equation 7.3.1 and Equation 7.3.2 are balanced~~

~~chemical equations. What is different on each side of  
the equation is how the atoms are arranged to make~~

~~molecules or ions. A chemical reaction represents a  
change in the distribution of atoms but not in the~~

~~number of atoms. In this reaction, and in most  
chemical reactions, bonds are broken in the reactants~~

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(here C-O and N-H bonds), and new bonds are formed to create the products (here, O-H and N-N bonds).

Chapter 7.3: Chemical Equations - Chemistry  
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Chapter 7 3 Answers Chemical CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Label each of the following statements as True or False: True a. If the formula mass of one molecule is x amu, the molar mass is x g/mol. False b. Samples

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CHAPTER 7 REVIEW Chemical Formulas and Chemical  
Compounds SECTION 3 SHORT ANSWER Answer the  
following questions in the space provided. 1. Label  
each of the following statements as True or False:  
True a. If the formula mass of one molecule is  $x$  amu,  
the molar mass is  $x$  g/mol. False b. Samples of equal  
numbers of moles of two different chemicals

7 Chemical Formulas and Chemical Compounds  
Chapter 7 3 Answers Chemical Formulas And  
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wanted to give away for free. Chapter 7 3 Answers  
Chemical As shown in Figure 7.3.1, applying a small  
amount of heat to a pile of orange ammonium  
dichromate powder results in a vigorous reaction  
known as the Page 4/30

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ethanol (CH<sub>3</sub>CH<sub>2</sub>OH) Answer: Ba, +2; F, -1; C, 0;  
H, +1; O, -2; K, +1; Cr, +6; O, -2; Cs, +1; O, -½; C,  
-3; H, +1; C, -1; H, +1; O, -2; H, +1

## Chapter 7.5: Types of Chemical Reactions - Chemistry

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Answer. (a) baking soda (sodium bicarbonate) plus  
vinegar (acetic acid) yields carbon dioxide plus water  
plus sodium ion plus acetate ion. Vinegar (Acetic acid)  
+ Baking soda (Sodium hydrogencarbonate) ->  
Carbon dioxide + other substances. N

$\text{NaHCO}_3 + \text{CH}_3\text{COOH} \rightarrow \text{NaC}_2\text{H}_3\text{O}_2 + \text{H}_2\text{O} + \text{CO}_2$   
 $\text{O}_3 + \text{C H}_3 \text{C } 00 \text{ H} - > \text{N a C } 2 \text{ H } 3 \text{ O } 2 + \text{H } 2 \text{ O} + \text{C}$   
 $\text{O } 2.$

## Physical and chemical changes class 7 Questions with Answers

Chapter 7: Chemical Nomenclature 7.1 Molecular  
Formula Practice Questions ... 7.3 Cations Practice ...  
3) 2 d. NaCN Answers 1. A compound with three or  
more elements. 2. Same rule as naming binary ionic  
compounds. 3. a. Hg(NO

CK-12 Chemistry Concepts - Intermediate Answer Key  
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## Chapter 7 Chemical Reactions Answer Key

7.E: Chapter 7 Homework Answers Last updated; Save as PDF Page ID 202166; No headers. Evidence of a Chemical Reaction. 1. a) b) 3. c) Writing and Balancing Chemical Equations. 5. b) 7. a)  $\text{Zn(s)} + 2\text{HCl(aq)} \rightarrow \text{ZnCl}_2(\text{aq}) + \text{H}_2(\text{g})$  b)  $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O(l)}$  c)  $\text{Ca(OH)}_2(\text{aq}) + \text{CO}_2(\text{g}) \rightarrow \text{CaCO}_3(\text{s}) + \text{H}_2\text{O(l)}$  d)  $2\text{Fe}_2\text{O}_3 \dots$

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Iron and sulfur react to form Iron (III) sulfide 2. Phosphorus and oxygen combine to form diphosphorus pentoxide 3. Water decomposes to form hydrogen and oxygen 4. Sodium reacts with water to form sodium hydroxide and hydrogen 5. Iron metal reacts with copper (II) nitrate to form iron (II) nitrate and copper metal 6.

Class Worksheet-2 Chapter 7 Name More Practice Pre  
...

Balancing Chemical Equations Chapter 7 Balancing

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Chemical Equations Answer Key Balance the  
equations below: 1)  $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$  2)  $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$  3)  $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$   
4)  $2 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$  5)  $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2$

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