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Polar \u0026amp; Non-Polar Molecules: Crash Course Chemistry #23 **12. The Shapes of Molecules: VSEPR Theory Polar Bonds, Polarity and Intermolecular Forces** **CHEMISTRY DK014 - TOPIC 7.2 (Part 1) - VSEPR Theory and 5 Basic Molecular Geometry** [Vsepr And Imf Homework](#)

VSEPR and IMF Homework Problems Lewis Structure Practice #1 Draw the Lewis structure for each of the following molecules: 1. Br 2. Methane 3. H₂S 4. Nitrogen trichloride 5. Phosphorus pentabromide 6. Boron trihydride 7. Carbon disulfide Lewis Structure Practice #2 Draw the Lewis structure for each of the following molecules 1. Arsenic triiodide 2.

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Vsepr And Imf Homework The valence shell electron pair repulsion (VSEPR) theory (or "VESPER" for short) is how the geometry of a molecule is determined around a central atom. The molecular geometry main shapes are tetrahedral, trigonal planar, trigonal pyramidal, bent, and linear and are named by measuring Page 3/11 Vsepr And Imf Homework -

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